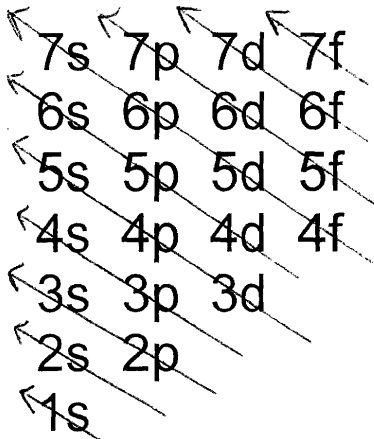


Electron Configuration "Cheat Sheet!"

e- capacities for sublevels



<i>s</i> or 0	<i>p</i> or 1	<i>d</i> or 2	<i>f</i> or 3
2e-	6e-	10e-	14e-

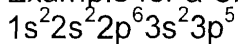
(1 pair e⁻) (3 pair e⁻) (5 pair e⁻) (7 pair e⁻)

Nucleus

Know how to do the following:

1. Electron Configurations for atoms:

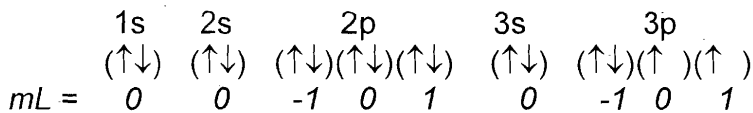
Example for a Chlorine atom (17 electrons):



(Note that the superscripts add up to 17)

2. Hund's Orbital Diagrams (arrow diagrams):

Example for Sulfur (16 electrons) Each parenthesis represents an **Orbital**.



3. Quantum Number Summary for an atom:

Example: All 5 electrons in Boron:

	1 st e-	2 nd e-	3 rd e-	4 th e-	5 th e-
1 st Q# n	1	1	2	2	2
2 nd Q# L	0	0	0	0	1
3 rd Q# m _L	0	0	0	0	-1
4 th Q# m _s	½	-½	½	-½	½